Name: Date: Pd:

ATOMIC BASICS PRACTICE

Parts of the Atom:

SUBATOMIC PARTICLE	ELECTRIC CHARGE	LOCATION IN ATOM

Complete the table for the elements:

ELEMENT NAME	ATOMIC NUMBER	AVERAGE ATOMIC MASS	PROTONS	ELECTRONS
Hydrogen				
Boron				
Nitrogen				
Oxygen				
Neon				

For each of the following ions, indicate the total number of protons and electrons in the ion:

Ion	Number of Protons	Number of Electrons
Cl^{-1}		
K^{+1}		
S^{-2}		
Sr^{+2}		
Al^{+3}		
P^{-3}		

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Here are three isotopes of an element: $_{6}^{12}C$ $_{6}^{13}C$ $_{6}^{14}C$

a. The element is: _____

b. The number 6 refers to the _____

c. The numbers 12, 13, and 14 refer to the _____

d. How many protons and neutrons are in the first isotope? _____

e. How many protons and neutrons are in the second isotope? _____

f. How many protons and neutrons are in the third isotope? _____

Complete the following chart:

Isotope name	atomic #	mass #	# of protons	# of neutrons	# of electrons	Isotopic Symbol
uranium-235						
uranium-238						
boron-10						
boron-11						

Fill in the following chart:

Element/Ion	Atomic Number	Number of Protons	Number of Neutrons	Number of Electrons	Mass Number
$_{1}^{1}H$					
$_{1}^{1}H^{+}$					
$_{3}^{7}Li$					
$_{17}^{35}Cl^{-}$					
$_{12}^{24}Mg^{2+}$					
$_{33}^{75}As$					
$_{47}^{108}Ag^{+}$					
$_{16}^{32}S^{2-}$					
		30		28	66
	76		114		

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Complete the following diagram: