Name: _____ Date: _____ Pd: _____

Electron Configurations

Longhand Electron Configuration

Use the patterns within the periodic table to write longhand electron configurations for the following atoms.

	Symbol	# e^-	Longhand Electron Configuration
1.	Mg		
2.	P		
3.	V		
4.	Ge		
5.	Kr		
6.	U		

Name: _____ Date: _____ Pd: _____

HONORS - ELECTRON CONFIGURATION PRACTICE

In the space below, write the full (unabbreviated) electron configurations of the following elements:

1) Na _____

2) Fe _____

3) Br^{-1} _____

4) Ba^{+2} _____

5) Np _____

In the space below, write the Noble Gas (abbreviated) electron configurations of the following elements:

6) Co _____

7) Ag _____

8) Te _____

9) Fr^{+1} _____

10) Cl^{-1} _____

Determine what elements are denoted by the following electron configurations:

11) $1s^2 2s^2 2p^6 3s^2 3p^4$ _____

12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ _____

13) [Kr] $5s^2 4d^{10} 5p^3$ _____

14) [Xe] $6s^2 4f^{14} 5d^6$ _____

15) [Rn] $7s^2 5f^{11}$ _____

State whether the following electron configurations is valid or invalid. If invalid, correct it!

16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ _____

17) $1s^2 2s^2 2p^6 3s^3 3d^5$ _____

18) [Ra] $7s^2 5f^8$ _____

19) [Kr] $5s^2 4d^{10} 5p^5$ _____

20) [Xe] _____