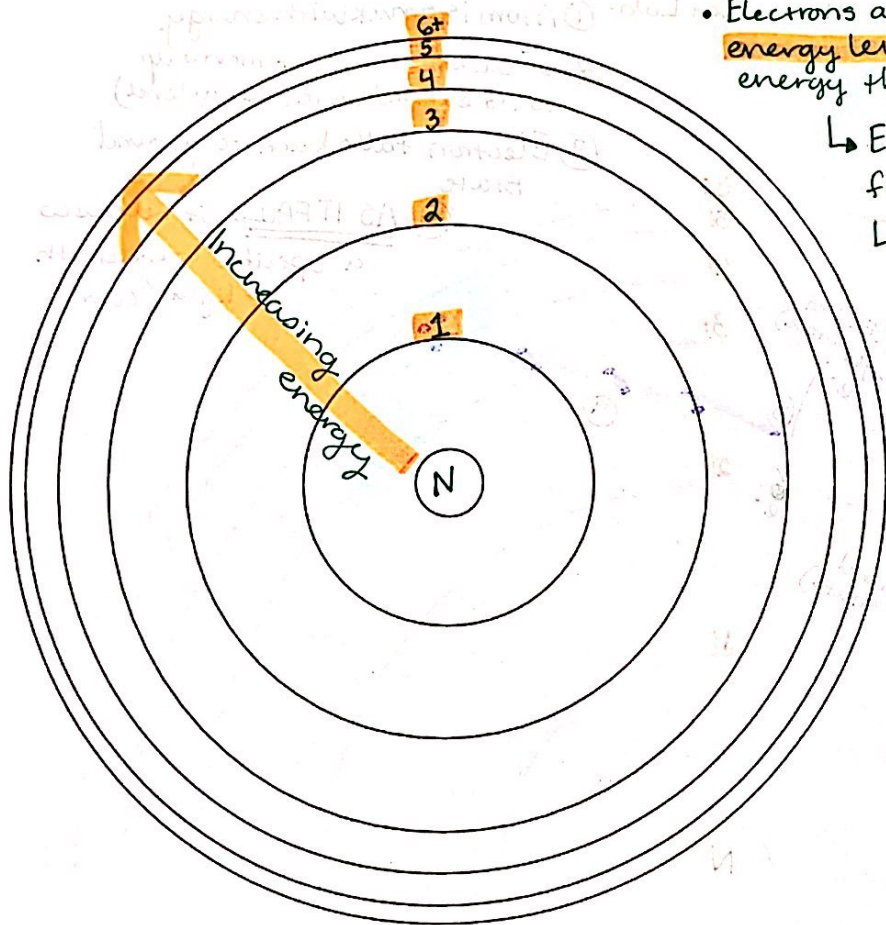


Name:

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ORGANIZATION OF THE ATOM - BOHR MODEL



- Electrons are arranged around the atom in energy levels (zones) based on how much energy they have.

↳ Energy increases as you move outward from the nucleus.

↳ 1st energy level - close to nucleus; lowest energy; holds 2e⁻

↳ 2nd energy level + beyond - up to 7; increase in #e⁻ but they still only hold max 8 valence e⁻

• Electron Organization:

Atom

↳ Energy Levels *

↳ Sublevels

↳ Orbitals (e⁻)

Name:

Date:

Pd:

ABSORPTION AND EMISSION OF ENERGY FROM ATOMS

- Electrons in energy levels can gain/lose energy and move up/down energy levels.

* When they move down, they release energy as a wavelength of light

↳ Every element releases a specific set of wavelengths

Flame Test Lab

- ① Atom is struck with energy
- ② e^- absorbs energy + moves up to its excited state (energy level)
- ③ Electron falls back to ground state
- ④ AS IT FALLS, it releases a specific wavelength of light (color)

