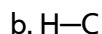
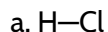


7. The bonds between the following pairs of elements are covalent. Arrange them according to polarity, naming the most polar bond first.



8. Using the table of electronegativities from your Periodic table, **calculate the EN difference** for the atoms that are bonded in the following molecules. Then tell whether the bond is **nonpolar covalent, polar covalent, or ionic**. Tell which atom has the **greater share of the bonding electrons**. In your drawing, show **which atom is partially positive or partially negative and draw the bond dipoles**.

<u>Molecule</u>	<u>EN Difference</u>	<u>Type of Bond</u>	<u>Atom with greater EN</u>	<u>Electron Dot Structure (Rough Draft)</u> <u>Add Bond Dipoles</u>
PCl ₃				
NH ₃				
H ₂ O				
Br ₂				
NI ₃				

What questions do you still have about bond polarity?